

# Gravimetric Analysis Calculation Questions

Gravimetric Analysis Calculation Questions Gravimetric Analysis Calculation Questions A Comprehensive Guide Gravimetric analysis is a fundamental analytical technique used to determine the mass of a specific analyte in a sample This method relies on the precise measurement of the mass of a precipitate formed through a chemical reaction between the analyte and a suitable reagent The principle of gravimetric analysis is straightforward the mass of the precipitate is directly proportional to the mass of the analyte in the original sample This guide provides a comprehensive overview of gravimetric analysis calculation questions covering various aspects of the technique and offering a stepbystep approach to solving problems Key Concepts and Definitions Before diving into the calculation questions its crucial to understand the key concepts and definitions associated with gravimetric analysis Analyte The specific component of interest in a sample that is being analyzed Precipitate The solid compound formed during the chemical reaction between the analyte and the reagent Precipitating Agent The reagent used to react with the analyte and form the precipitate Gravimetric Factor The ratio of the molecular weight of the analyte to the molecular weight of the precipitate This factor allows for the conversion of the mass of the precipitate to the mass of the analyte Stoichiometry The study of the quantitative relationships between reactants and products in chemical reactions It helps in understanding the mole ratios involved in the reaction Types of Gravimetric Analysis Gravimetric analysis can be broadly classified into two main categories 1 Precipitation Gravimetry This method involves the precipitation of the analyte as an insoluble compound by adding a suitable precipitating agent The precipitate is then filtered washed dried and weighed 2 Volatilization Gravimetry In this method the analyte is selectively volatilized from the 2 sample by heat or chemical reaction The volatile component is collected and measured and the mass of the analyte is determined by difference Common Calculation Questions Lets delve into some common calculation questions encountered in gravimetric analysis 1 Determining the Mass of Analyte from the Mass of Precipitate Problem A sample of limestone  $\text{CaCO}_3$  weighing 0500 g was dissolved in acid and the calcium ions were precipitated as calcium oxalate  $\text{CaC}_2\text{O}_4$  The mass of the dried calcium oxalate precipitate was found to be 0350 g Calculate the percentage of calcium carbonate in the limestone sample Solution 1 Write the balanced chemical equation for the reaction  $\text{CaCO}_3 + \text{H}_2\text{SO}_4 \text{aq} \rightarrow \text{CaSO}_4 \text{aq} + \text{H}_2\text{O} + \text{CO}_2 \text{g}$   $\text{CaSO}_4 \text{aq} + \text{Na}_2\text{C}_2\text{O}_4 \text{aq} \rightarrow \text{CaC}_2\text{O}_4 \text{s} + \text{Na}_2\text{SO}_4 \text{aq}$  2 Calculate the gravimetric factor Gravimetric Factor  $\frac{\text{Molecular Weight of CaCO}_3}{\text{Molecular Weight of CaC}_2\text{O}_4}$  Gravimetric Factor  $\frac{10009 \text{ g/mol}}{12810 \text{ g/mol}}$  0782 3 Calculate the mass of  $\text{CaCO}_3$  in the precipitate Mass of  $\text{CaCO}_3$   $\frac{\text{Mass of CaC}_2\text{O}_4}{\text{Gravimetric Factor}}$  Mass of  $\text{CaCO}_3$   $\frac{0350 \text{ g}}{0782}$  0274 g 4 Calculate the percentage of  $\text{CaCO}_3$  in the limestone sample  $\frac{\text{CaCO}_3 \text{ Mass of CaCO}_3}{\text{Limestone Sample}} \times 100$   $\frac{0274 \text{ g}}{0500 \text{ g}} \times 100$  548 2 Calculating the Mass of Precipitate from the Mass of Analyte Problem A solution containing 0100 g of chloride ions  $\text{Cl}^-$  is treated with silver nitrate solution to precipitate silver chloride  $\text{AgCl}$  Calculate the theoretical yield of  $\text{AgCl}$  precipitate Solution 3 1 Write the balanced chemical equation for the reaction  $\text{AgNO}_3 \text{aq} + \text{Cl}^- \text{aq} \rightarrow \text{AgCl} \text{s} + \text{NO}_3^- \text{aq}$  2 Calculate the moles of chloride ions Moles of  $\text{Cl}^-$   $\frac{\text{Mass of Cl}^-}{\text{Molecular Weight of Cl}^-}$  Moles of  $\text{Cl}^-$   $\frac{0100 \text{ g}}{3545 \text{ g/mol}}$  000282 mol 3 Use the stoichiometric ratio from the balanced equation to determine the moles of  $\text{AgCl}$  Moles of  $\text{AgCl}$  Moles of  $\text{Cl}^-$

Moles of AgCl 0.00282 mol 4 Calculate the mass of AgCl precipitate Mass of AgCl  
 Moles of AgCl x Molecular Weight of AgCl Mass of AgCl 0.00282 mol x 143.32 g/mol  
 0.404 g 3 Calculating the Concentration of Analyte Problem A 1000 mL sample of  
 a solution containing an unknown concentration of barium ions  $Ba^{2+}$  was treated  
 with excess sodium sulfate  $Na_2SO_4$  to precipitate barium sulfate  $BaSO_4$ . The  
 precipitate was filtered, dried, and weighed to be 0.150 g. Calculate the  
 concentration of barium ions in the original solution in mg/L. Solution 1 Calculate the  
 moles of  $BaSO_4$ . Moles of  $BaSO_4$  Mass of  $BaSO_4$  Molecular Weight of  $BaSO_4$   
 Moles of  $BaSO_4$  0.150 g / 233.39 g/mol = 0.000643 mol 2 Use the stoichiometric ratio  
 from the balanced equation to determine the moles of  $Ba^{2+}$ . Moles of  $Ba^{2+}$  Moles of  
 $BaSO_4$  Moles of  $Ba^{2+}$  0.000643 mol 3 Calculate the concentration of  $Ba^{2+}$  in the  
 original solution. Concentration of  $Ba^{2+}$  Moles of  $Ba^{2+}$  Volume of Solution x 1000  
 mg/g Concentration of  $Ba^{2+}$  0.000643 mol / 0.100 L x 1000 mg/g = 6.43 mg/L 4 Practical  
 Applications of Gravimetric Analysis Gravimetric analysis finds wide applications in  
 various fields including Environmental Chemistry Determining the levels of heavy  
 metals, pollutants, and other contaminants in water, soil, and air. Food Chemistry  
 Analyzing the composition of food products such as the fat content in milk or the  
 salt content in processed foods. Pharmaceutical Analysis Determining the purity and  
 potency of drugs and pharmaceuticals. Industrial Chemistry Monitoring and  
 controlling the quality of raw materials and finished products in various industries.  
 Forensic Science Analyzing evidence to identify substances and trace elements.  
 Challenges and Limitations While gravimetric analysis is a powerful technique, it  
 also has its limitations. Time-consuming The process can be time-consuming due to  
 the need for filtration, washing, drying, and weighing steps. Not suitable for all  
 analytes Some analytes may not form suitable precipitates or may be volatile at  
 high temperatures. Sensitivity The method may not be sensitive enough to detect  
 trace levels of analytes. Interferences Other components in the sample may  
 interfere with the precipitation reaction, leading to inaccurate results. Conclusion  
 Gravimetric analysis is a versatile analytical technique that provides accurate and  
 precise results when properly executed. Understanding the key concepts, mastering  
 the calculation procedures, and recognizing potential challenges are crucial for  
 successful application of this method. By carefully selecting the appropriate  
 precipitation reaction and controlling the experimental conditions, gravimetric  
 analysis remains a valuable tool for determining the mass of analytes in various  
 samples across diverse fields.

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